

## Electrolyte Solutions Robinson Stokes

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### Electrolyte Solutions Robinson Stokes

Electrolyte Solutions, von R. A. Robinson und R. H. Stokes. — The Measurement and Interpretation of Conductance, Chemical Potential and Diffusion in Solutions of Simple Electrolytes. Butterworths Scientific Publication, London 1959. 2.

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Electrolyte solutions. R. A. Robinson and R. H. Stokes. Academic Press, New York, 1959. xv + 559 pp. \$11.50.

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In very dilute regions, the Stokes and Robinson's equation reduces to a much simpler form, i. e. Activity coefficients of an electrolyte at lower concentrations, say up to 0.1 m, can be calculated from the equation provided suitable values of  $\gamma_{\pm}$ ; and  $h$  are available.

### A simplified form of Stokes-Robinson equation - Canadian ...

The proposed extension to the Stokes and Robinson hydration theory is based on the premise that H<sub>2</sub>O molecules of electrolyte solutions reside in two unique environments, the solvent environment and the hydration envelope (or sheath) environment.

### The stokes and robinson hydration theory: A modification ...

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Chemical Engineering Thermodynamics, Newman, S.A., Ed., Ann Arbor Science, p. 37990 Robinson, R.A. and Stokes, R.H., 1965. Electrolyte Solutions, Second edition (revised)

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### **Electrolyte solutions - ScienceDirect**

Robinson RA, Stokes RH, Electrolyte Solutions. Butterworths: London, 1955, 30. Studies on solute-solvent interactions of amino acids in water + D-glucose mixtures at different temperatures

### **Thermo-Acoustic Parameters of Multi-Charged Electrolytes ...**

Robinson, R. A., Stokes, R. H. This classic text deals primarily with measurement, interpretation of conductance, chemical potential, and diffusion in electrolyte solutions. The book provides detailed theoretical interpretations, plus extensive tables of thermodynamic and transport properties.

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### **Electrolyte Solutions by Robinson R a Stokes R H - AbeBooks**

Precise vapor pressure data for LiBr solutions in 2-propanol are given for temperatures ranging from (298.15 to 348.15) K. The molality range was from (0.0743 to 1.4825) mol·kg<sup>-1</sup>.

### **Data Banks of Electrolytes - ResearchGate**

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ROBINSON R. H. and STOKES. Academic Press, New York, 1959. xv 559 pp. \$11.50. Those chemists who found in the first edition of *Electrolyte Solutions* both a lucid and elegant discussion of the measurement and interpretation of conductance, chemical potential, diffusion and related phenomena in electrolytic solutions will welcome this second edition.

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E.g., for a 1-molal solution of 1-1 electrolyte at 25 °C,  $\kappa = 0.03 \text{ nm}^{-1}$  When  $m = 0.001 \text{ mol kg}^{-1}$ ,  $\kappa = 9.64 \text{ nm}^{-1}$ . Electrolyte solutions are nonideal even at low concentration because of long-range Coulombic forces. From the Debye-Hückel theory, the activity coefficient  $\gamma_i$  of an ion with charge  $z_i$  in a dilute solution of ionic strength  $I$  is

### **CH. 9 Electrolyte Solutions**

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