

Chapter 18 Review Chemical Equilibrium

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CHAPTER 18: Chemical Equilibrium I. a. Write the chemical equilibrium expression for the following equations. Include the value of K. (a) (g) (g) $K=0.1$ (b) (g) +02 (g) (g) $K = 1.8 \times 10^{-}$ Does the reaction favor products or reactants? (Will there be mostly products or reactants when it reaches 2. a. Compare the rates of forward and reverse reactions when equilibrium has been reached. b.

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Modern Chemistry 145 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. ____ Raising the temperature of any equilibrium system always (a) favors the forward reaction. (b) favors the reverse reaction. (c) favors the exothermic reaction. (d) favors the endothermic reaction. 2. Consider the following equilibrium equation:

CHAPTER 18 REVIEW Chemical Equilibrium

Chemical equilibrium between a saturated solution of ions and remaining solid. Precipitation. Process of aqueous ions forming a solid (1 process in equilibrium) Dissociation. Process of an ionic solid separating into aqueous ions (1 process in equilibrium) Until they are in equilibrium with their ions, ionic solids will.

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Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the K_{sp} for that solution, listed on the left. ____ The ion product exceeds the K_{sp} .

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CHEMICAL EQUILIBRIUM 553 Copyright © by Holt, Rinehart and Winston. All rights reserved. SECTION 18-1 OBJECTIVES Define chemical equilibrium. Explain the nature of the equilibrium constant. Write chemical equilibrium expressions and carry out calculations involving them. FIGURE 18-1 When heated, mercury(II) oxide decomposes into the elements from which it was

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Chemistry: Chapter 18- Chemical Equilibrium. Equilibrium. Value of K_c . $K_c < 1$. $K_c > 1$. when the rates of opposing actions are the same. represents the reaction progression... -all K_c 's are dependent o... reactants of the reaction are favored. products of the reaction are favored.

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Chapter 18 Review Chemical Equilibrium Answers Section 1

Chapter 18 Review "Reaction Rates and Equilibrium" Name: _____ 1. Energy that is available to do work is called free energy. 2. Reaction rate is defined as the number of atoms, ions, or molecules that react in a given time to form products. 3.

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chapter 18 chemical equilibrium modern Flashcards. A chemical reaction in which the products can react to reform.... When the rate of its forward reaction equals the rate of its r.... The ratio of the mathematical product of the concentrations of.... A chemical reaction in which the products can react to reform....

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chapter 18 review chemical equilibrium CHAPTER 18: Chemical Equilibrium I. a. Write the chemical equilibrium expression for the following equations. Include the value of K. (a) (g) $K=0.1$ (b) (g) $+O_2$ (g) (g) $K = 1.8 \times 10^{-2}$ Does the reaction favor products or reactants? (Will there be mostly products or reactants when it reaches 2. a.

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